

# Ab Initio Calculation of Vibrational Circular Dichroism Spectra Using Accurate Post-Self-Consistent-Field Force Fields: trans-2,3-Dideuteriooxirane

P. J. Stephens
F. J. Devlin
UNIVERSITY OF SOUTHERN CALIFORNIA

K. J. Jalkanen UNIVERSITY OF ILLINOIS

C. F. Chabalowski U.S. ARMY RESEARCH LABORATORY

ARL-TR-963

February 1996

19960213 068

DTIC QUALITY INSPECTED 4

#### **NOTICES**

Destroy this report when it is no longer needed. DO NOT return it to the originator.

Additional copies of this report may be obtained from the National Technical Information Service, U.S. Department of Commerce, 5285 Port Royal Road, Springfield, VA 22161.

The findings of this report are not to be construed as an official Department of the Army position, unless so designated by other authorized documents.

The use of trade names or manufacturers' names in this report does not constitute indorsement of any commercial product.

### **REPORT DOCUMENTATION PAGE**

Form Approved
OMB No. 0704-0188

Public reporting burden for this collection of information is estimated to average 1 hour per response, including the time for reviewing instructions, searching existing data sources, gathering and maintaining the data needed, and completing and reviewing the collection of information. Send comments regarding this burden estimate or any other aspect of this collection of information, including suggestions for reducing this burden. to Washington Headquarters Services, Directorate for Information Operations and Reports, 1215 Jefferson Davis Highway, Suite 1204, Arlington, VA 22202-4302, and to the Office of Management and Budget, Paperwork Reduction Project (0704-0188), Washington, DC 20503.

1. AGENCY USE ONLY (Leave blank) 2. REPORT DATE 3. REPORT TYPE AND DATES COVERED							
		February 1996	Final, Januar	y 1992–January 1993			
4. TITLE AND SUBTITLE					ING NUMBERS		
41 7 13 64 4 3 65		-1 Dist -1 C	- TT-! A	DD.	1L161102AH43		
Ab Initio Calculation of Vibrat	PK:	ILIUIIVANINJ					
	Post-Self-Consistent-Field Force Fields: trans-2,3-Dideuteriooxirane						
6. AUTHOR(S)					·		
P. J. Stephens, K. J. Jalkanen,							
P. J. Stephens, R. J. Jankaren,							
7. PERFORMING ORGANIZATION N	8. PERF	ORMING ORGANIZATION					
		RT NUMBER					
U.S. Army Research Laborator	ry			A	ARL-TR-963		
ATTN: AMSRL-WT-PD	•						
Aberdeen Proving Ground, MI	D 21005-	-5066					
9. SPONSORING/MONITORING AG	ENCY NAM	ME(S) AND ADDRESS(ES	)		ISORING/MONITORING NCY REPORT NUMBER		
•							
11. SUPPLEMENTARY NOTES		<del></del>					
*Correspondence should be ad	dressed to	Mr. Stephens who, at	long with Mr. Devlin, v	vorks for	the University of Southern		
California. Mr. Jalkanen worl	ks for the	University of Illinois	and Mr. Chabalowski	works fo	r the U.S. Army Research		
Laboratory.							
12a. DISTRIBUTION / AVAILABILITY	12b. DIS	TRIBUTION CODE					
Approved for public release; d							
13. ABSTRACT (Maximum 200 words)							
The ab initio calculation of harmonic force fields using post-self-consistent-field (post-SCF) methods has recently							
increased greatly in efficiency. As a result, accurate post-SCF harmonic force fields are now achievable for experimentally							
accessible chiral molecules. This in turn, permits for the first time, comparison to experiment of ab initio predictions of vibrational circular dichroism (VCD) spectra based on accurate harmonic force fields. This advance is exploited here in							
vibrational circular dichroism (VCD) spectra based on accurate harmonic force fields. This advance is exploited here in the calculation of the VCD spectrum of trans-2,3-dideuteriooxirane using a large basis set MP2 force field. Excellent							
accord with recent experimental work is achieved. With the exception of C-H and C-D stretching modes, residual							
differences between predicted and experimental VCD intensities can be attributed to the incomplete inclusion of correlation							
in the calculation of atomic axial tensors.							
14. SUBJECT TERMS					15. NUMBER OF PAGES 21		
vibrational circular dichroism s	•		m chemical calculations	S	16. PRICE CODE		
of molecular spectra, dichroisr	m, vibrati	onal spectra					
17. SECURITY CLASSIFICATION		RITY CLASSIFICATION	19. SECURITY CLASSIFI	CATION	20. LIMITATION OF ABSTRACT		
OF REPORT	_	HIS PAGE	OF ABSTRACT		тπ.		

INTENTIONALLY LEFT BLANK.

### **ACKNOWLEGMENTS**

P. J. Stephens is grateful to the National Science Foundation (NSF), the National Institutes of Health (NIH), the North American Treaty Organization (NATO), and the San Diego Supercomputer Center for support of work on vibrational circular dichroism at the University of Southern California (USC) over the years. C. F. Chabalowski is grateful to the U.S. Army Edgewood, Research, Development, and Engineering Center (ERDEC) for partial support of this work.

INTENTIONALLY LEFT BLANK.

### TABLE OF CONTENTS

		<u>Page</u>
	ACKNOWLEGMENTS	iii
1.	INTRODUCTION	1
2.	RESULTS AND DISCUSSION	1
3.	CONCLUSION	8
4.	REFERENCES	10
	DISTRIBUTION LIST	13

INTENTIONALLY LEFT BLANK.

#### 1. INTRODUCTION

The prediction of the vibrational circular dichroism (VCD) spectrum of a chiral molecule at the harmonic level of approximation requires the calculation of its equilibrium geometry, harmonic force field, atomic polar tensors, and atomic axial tensors [1–3]. These molecular properties are most accurately calculated using ab initio methods [4–7]. While predictions of VCD spectra based on ab initio calculations at the self-consistent-field (SCF) level of approximation have achieved substantial success in reproducing experimental spectra [8–25], they have been significantly limited by the inherent deficiencies of the SCF approximation. Accurate calculations require post-SCF methods.

Recently, the feasibility of the calculation of accurate harmonic force fields using post-SCF methods has increased sufficiently to permit their use in predicting VCD spectra of experimentally accessible chiral molecules. Consequently, it is now practicable to compare experimental VCD spectra to theoretical predictions which are not limited in accuracy by uncertainties in the harmonic force field. In turn, the origins of the remaining discrepancies between theory and experiment can be more precisely analyzed. Here we report the first calculations of VCD spectra which take advantage of this progress. Specifically, we report calculations of the VCD spectrum of trans-2,3-dideuteriooxirane (1) based on a harmonic force field obtained at the MP2 level of approximation using a large basis set. The results are compared to the recently published experimental VCD spectra of Freedman et al. [26] for solutions of 1. The accuracy of the calculations is limited only by the incomplete inclusion of correlation in the calculation of atomic axial tensors and by the absence of any consideration of anharmonicity and condensed-phase effects. The magnitudes of the errors arising from the neglect of the contributions are evaluated. The analysis is assisted by comparisons of experimental vibrational frequencies and unpolarized absorption intensities for 1, oxirane 2, and oxirane- $d_4$  (3)—in solution for 1 [26] and in the gas phase for 2 and 3 [27–28]—to the results of calculations using the same force field and atomic polar tensors as used in predicting the VCD of 1. There results a more sophisticated analysis of the deviations between a calculated and an experimental VCD spectrum than heretofore achievable.

### 2. RESULTS AND DISCUSSION

The harmonic vibrational frequencies and rotational strengths of 1, calculated using a [8s6p3d/6s3p] basis set (hereafter referred to as VD/3P) [7], are given in Table 1. The harmonic force field and atomic polar tensors  $P_{\alpha\beta}^{\lambda}$  were obtained at the MP2 level via analytical derivative techniques [29] using a

Table 1. Vibrational Frequencies, Intensities, Dipole Strengths, and Rotational Strengths of 1<sup>a</sup>

									-						1		
	APŢ	ဦ ဝ	-6.3	9.0	-1.0	2.9	1.6	9.0	-1.0	-7.5	0.5	8.7	-2.0	6.0	0.2	-2.6	9.0-
	AAT	ဗိ							13.2	1.4	1.0	-0.8	-2.5				
	exb <sub>p</sub>		11.4	6.8-	12.1	-10.4	(-15)	(-2.5)	24.1	-4.9	11.1	-29.0	-6.2	(~ + 5)	<b>(</b> +)		
R		(P·M/P·L)	(18.6/4.2)	(-20.4/-4.0)	(9.3/3.3)	(-8.1/-3.1)	(-8.1/-0.2)	(-1.2/-0.1)	(12.4/–1.5)	(-5.4/-0.8)	(2.3/7.8)	(-30.5/2.3)	(0.9/–4.6)	(17.0/–9.0)	(0.3/1.4)	(6.4/5.3)	(0.1/-0.2)
	Calc	MP2f,8	22.8	-24.4	12.7	-11.2	-8.3	-1.3	10.9	-6.3	10.1	-28.2	-3.7	8.0	1.7	11.8	0.0
		SCF	29.1	-25.0	13.7	-14.1	6.6-	-1.9	11.9	1.2	9.6	-36.9	-1.7	7.1	1.5	14.4	9.0
	exb			53	5.7	27	(12)	(2.3)	29.6		8.6	54	6.3				
Q		MP2f	31.0	7.4	9.8	27.7	8.4	9.0	19.7	1.9	6.3	57.1	11.9	122.9	38.7	113.9	8.0
	Calc	SCF	50.9	7.4	9.0	45.9	12.6	1.1	26.7	0.1	11.1	108.3	20.1	192.0	59.3	148.2	1.5
I	exb		,	40.1	3.2	15.1			9.1		2.4	12.8	1.4				
	Calc	MP2f	25.0	6.0	5.1	16.4	3.0	0.2	6.3	0.5	1.8	13.9	2.8	27.7	8.0	22.1	0.1
	qdxə		3,027	3,014	2,252	2,232			1,226	1,109	1,102	948	914				
ų	ြင	В	3,217			2,353		1,380			1,128		922		820		675
	Calc	А		3,212	2,366		1,439		1,266	1,147		973		006		774	

\* Frequencies  $\bar{v}$ , in cm<sup>-1</sup>; dipole strengths, D, in  $10^{-40}$  esu<sup>2</sup> cm<sup>2</sup>; intensities, I, in km/mol; rotational strengths, R, in  $10^{-44}$  esu<sup>2</sup> cm<sup>2</sup>. Rotational strengths are for the (5,5)

enautiomer.

<sup>b</sup> From Freedman et al. [26], in C<sub>2</sub>Cl<sub>4</sub> solution for C-H and C-D stretching modes and in CS<sub>2</sub> solution otherwise; values in parentheses estimated from gas-phase data. Rotational strengths are not normalized to 100% enantiomeric excess (ee); ee for (S,S)-1 and (R,R)-1 were estimated to be 94%.

<sup>c</sup> Contribution of correlation correction to local atomic axial tensors to rotational strength: R(exp) – R(MP2).

<sup>d</sup> Contribution of correlation correction to atomic polar tensors to rotational strength: R(MP2)–R(SCF).

<sup>e</sup> Calculation using SCF atomic polar tensors.

<sup>f</sup> Calculated using MP2 atomic polar tensors.

CRAY-2 version of GAUSSIAN 92 [30]. Atomic axial tensors,  $M_{\alpha\beta}^{\lambda}$ , were calculated using the distributed origin gauge [2,7,13,31,32], when

$$M_{\alpha\beta}^{\lambda} = (I_{\alpha\beta}^{\lambda})^{\lambda} + \frac{i}{4\hbar c} \sum_{\gamma\delta} \varepsilon_{\beta\gamma\delta} R_{\lambda\gamma}^{0} P_{\alpha\beta}^{\lambda} , \qquad (1)$$

where  $(I_{\alpha\beta}^{\lambda})^{\lambda}$  is the atomic axial tensor of atom  $\lambda$  evaluated with the origin at the equilibrium position of nucleus  $\lambda$ ,  $\overrightarrow{R}_{\lambda}^{0}$ . The "local" atomic axial tensors,  $(I_{\alpha\beta}^{\lambda})^{\lambda}$ , were calculated at the SCF level via analytical derivative techniques [33,34] using a Cray Y-MP version of CADPAC 5.0 [35]. The atomic axial tensors are thus only partially corrected for correlation. The use of the distributed origin gauge guarantees origin-independent rotational strengths [2]. The VCD spectrum derived from the calculated rotational strengths and the experimental VCD spectrum of Freedman et al. [26] are shown in Figures 1 and 2 and are in obvious correspondence, permitting unambiguous assignment of all experimental features to fundamental transitions. The assignment, together with the rotational strengths obtained by Freedman et al. from the experimental VCD spectrum, is given in Table 1.

The accuracy of predicted rotational strength is a composite function of the accuracies of the harmonic force field, the atomic polar tensor and the atomic axial tensors, and, in addition, the magnitude of anharmonicity and condensed-phase contributions, which are not included in the theoretical formalism. The accuracies of the force field and atomic polar tensors and the significance of anharmonicity can be independently assessed by comparison of the frequencies and unpolarized absorption intensities predicted for the undeuterated and perdeuterated isotopomers of oxirane, oxirane (2) and oxirane - d<sub>4</sub> (3), with the accurate gas-phase results of Nakananga [27–28]. As seen from Table 2, excluding C-H and C-D stretching modes, calculated and observed frequencies differ by 2.2% on average; the maximum error is 3.9%. The errors for C-H and C-D stretching modes are substantially larger: 5–7% and 4–5%, respectively. Although harmonic frequencies for isotopomers of oxirane have not been derived from experimental data, the contributions of anharmonicity are known in many other small molecules [4]. By comparison with such data, it is clear that the calculated errors for 2 and 3 are comparable to anharmonicity contributions and that, with the respect of the force field, the harmonic limit has been effectively realized. In the case of intensities, excluding the C-H and C-D stretching modes, calculated

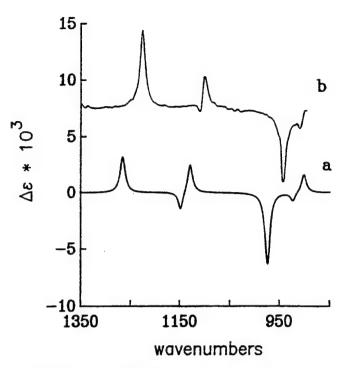


Figure 1. Predicted (a) and experimental (b) VCD spectra of 1. The predicted spectrum results from frequencies and rotational strengths are given in Table 1. Lorentzian band shapes are assumed [11,15]. The half-width at half-height (γ) is arbitrarily chosen to be 6.0 cm<sup>-1</sup> for all bands [26].

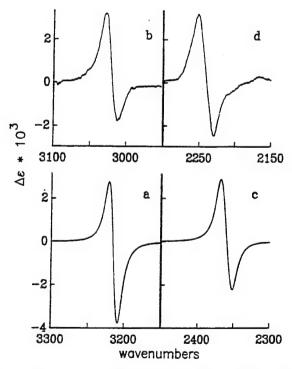


Figure 2. Predicted (a,c) and experimental (b,d) VCD spectra result from frequencies and rotational strengths given are in Table 1. Lorentzian band shapes are assumed [11,15]. The half-width at half-height (γ) is arbitrarily chosen to be 10.0 cm<sup>-1</sup> for all bands [26].

and experimental intensities differ on average by 0.8 km/mol, while the quoted experimental uncertainties are on average 0.9 km/mol. Thus, calculated and experimental intensities agree within experimental error. This result confirms the excellent accuracy of the force field and of the normal coordinates derived therefrom. In addition, it demonstrates that the errors in intensities arising from errors in the calculated atomic polar tensors and from anharmonicity are both less than experimental error. Thus, with respect to intensities the harmonic limit has been effectively reached and is sufficient. In the case of the C-H and C-D stretching modes, the differences between calculated and experimental intensities are significantly larger and, in addition, are substantially greater than the experimental errors. The differences for the C-H modes are appreciably larger than for the C-D modes. In view of the excellent accuracy of the calculations for all other modes, these discrepancies can only be due to anharmonicity. Thus, for these specific modes, the harmonic limit is clearly insufficient.

Since the same force field and atomic polar tensors determine the frequencies and intensities of all isotopomers of oxirane, the frequencies and intensities of 1 should be as accurately predicted as those of 2 and 3. Comparison to the results of Freedman et al. [26] (Table 1) supports this expectation. Excluding C-H and C-D stretching modes, the error in frequencies is 2.5% on average, the maximum error being 3.4%. The error in intensities is 1.5 km/mol on average. Although the number of modes on which these statistics are based is small, the average errors are very similar to those for 2 and 3. Experimental uncertainties were not reported for 1, and we cannot directly compare the differences between theory and experiment to experimental errors. In contrast to 2 and 3, the experimental parameters for 1 are derived from spectra in solution. The comparable accuracies of predicted frequencies and intensities for 1, 2, and 3 lead to the conclusion that condensed-phase effects are small, i.e., that they are less than either anharmonicity effects or experimental errors, whichever are limiting.

The predicted rotational strengths of 1 are in perfect agreement with the experimental parameters derived by Freedman et al. with regard to both sign and relative magnitude (Table 1). The differences in absolute magnitudes of predicted and experimental rotation strengths could originate either in theoretical or in experimental error. Experimental uncertainties were not reported and are not easily estimated. For three bands, rotational strengths are estimates from gas-phase data. In addition, the C-H and C-D stretching rotational strengths are particularly susceptible to uncertainty since they derive from the deconvolution of unresolved, oppositely signed couplets. To a substantial degree, the differences between predicted and experimental rotational strengths are likely to be attributable to experimental uncertainties. With regard to the calculations, leaving aside the C-H and C-D stretching modes, the previous discussion

Table 2. Vibrational Frequencies and Intensities for 2 and 3<sup>a</sup>

				Oxirane (2)	(2							ŏ	Oxirane-d <sub>4</sub> (3)	3)			
		ū					I				ū					I	
	S	calc			calc	ျ				ca	calc			ca	calc		
A <sub>1</sub>	A2	B <sub>1</sub>	B <sub>2</sub>	exp <sub>p</sub>	SCF <sup>c</sup>	MP2 <sup>d</sup>	6	exp <sup>b</sup>	A <sub>1</sub>	A <sub>2</sub>	$\mathbf{B}_1$	B <sub>2</sub>	exp <sub>p</sub>	SÇF	Mp2	ě	exp <sub>p</sub>
			3,269	3,065	40.3	25.3	36.24	(4.69)				2,43 9	2,322	28.6	18.0	23.77	(3.11)
	3,255				0	0				2,43				0	0		
3,164				3,005	13.4	13.4	6.04	(3.58)	2,308				2,204	9.2	8.7	7.17	(1.44)
		3,156		2,987 <sup>e</sup>	42.1	24.6	44.31	(5.95) <sup>e</sup>			2,276		2,177	25.2	14.7	19.89	(2.69)
1,555				1,497	2.2	1.5	1.04	(0.42)	1,346				1,312	9.4	6.0	6.43	(0.52)
		1,526		1,470	0.0	0.1	0.34	(0.25)			1,119		1,083	0.1	0.0	0.05	(0.11)
1,309				1,270	13.8	11.2	13.16	(0.45)	1,043				1,013	1.1	1.1	1.69	(0.78)
	1,182				0	0			086				896	46.4	22.8	23.14	(3.05)
			1,167		7.9	4.0				943				0	0		
		1,157			0.0	9.0	4.33	(0.48)			806		900	0.5	0.0	0.10	(0.14)
1,154					0.3	0.4						906	868	8.4	4.8	3.80	(0.57)
	1,061				0	0					811		812	11.1	7.5	6.76	(0.79)
890				877	101.2	8.99	63.94	(2.60)	770				752	48.8	37.6	34.23	(2.64)
			833	808	0.4	0.2	0.21	(0.21)		753				0	0		
		824		822	12.5	8.3	8.56	(080)				594	577	0.3	0.1	0.45	(0:30)

<sup>&</sup>lt;sup>a</sup> Frequencies,  $\bar{v}$ , in cm<sup>-1</sup>; intensities, I, in km/mol.
<sup>b</sup> Gas-phase frequencies and intensities from Nakanaga [27–28]; quoted errors are in parentheses.
<sup>c</sup> Calculated using SCF atomic polar tensors.
<sup>d</sup> Calculated using MP2 atomic polar tensors.
<sup>e</sup> Corrected for Fermi resonance with (1,497 + 1,470)-cm<sup>-1</sup> combination.

of unpolarized absorption intensities shows that errors in rotational strengths arising from errors in the harmonic force field and atomic polar tensors, from the neglect of anharmonicity and from the neglect of solvent effects, should be insignificant (assuming that solvent effects on VCD intensities are no longer than on unpolarized absorption intensities). There remain the errors arising from the incomplete inclusion of correlation in the atomic axial tensors—specifically, from the evaluation of the "local" atomic axial tensors ( $I_{\alpha\beta}^{\lambda}$ ) at the SCF level of approximation. If experimental errors are ignored, the magnitudes of these errors are the differences between calculated and experimental rotational strengths, given in Table 1. The uncertainties in these values are equal to the experimental uncertainties. In one case—the 1,266/1,226-cm<sup>-1</sup> band—the difference is large; in all other (four) cases it is relatively small. In the case of the C–H and C–D stretching modes, as discussed previously, anharmonicity is not negligible, and experimental rotational strengths are particularly uncertain. The considerable difference between the theoretical and experimental rotational strengths for the C–H stretching modes is not surprising, therefore, while the close agreement for the C–D stretching modes is likely to be fortuitous. For these modes, the contributions of errors in atomic axial tensors to differences between experimental and calculated rotational strengths, while unlikely to be negligible, cannot yet be disentangled.

The reasonableness of the preceding analysis can be further probed in several ways. First, we can examine the "P·M" and "P·L" contributions to the rotational strengths [2,15,36] originating respectively in the first and second terms in the atomic axial tensors (equation 1). The former is subject to the correlation errors of the SCF  $(I_{\alpha\beta}^{\lambda})^{\lambda}$  tensors while the second in not. Significant errors are thus to be expected only when the "P·M" term is dominant and not when the "P·L" term is dominant. The magnitudes of these two contributions are given in Table 1. In most cases, the "P·M" term is dominant. In the particular case of the 1,266/1,226-cm<sup>-1</sup> mode, the only band for which the estimated correlation error was large, the "P·M" contribution is overwhelmingly dominant. Our analysis is therefore not inconsistent with the relative magnitudes of "P·M" and "P·L" contributions to the rotational strengths.

In addition, we can examine the sensitivity of rotational strengths to the exclusion of correlation in the calculation of the atomic polar tensors. VD/3P atomic polar tensors were calculated at the SCF level of approximation (using analytical derivative methods and the Cray Y-MP implementation of CADPAC 5.0) and substituted for the MP2 atomic polar tensors in calculating unpolarized absorption intensities for 1, 2, and 3 and rotational strengths for 1, with the results given in Tables 1 and 2. With very few exceptions, unpolarized absorption intensities are increased, the largest increases being around a factor of 2. The rotational strengths of 1 are affected in a more complex manner, as is to be expected since changes

in the atomic polar tensors affect both electric and magnetic dipole transition moments, in magnitude and in direction. In one case, the 1,147/1,109-cm<sup>-1</sup> transition, the sign is changed. The magnitudes of the changes are given in Table 1 and can be seen to vary considerably. The errors in the calculated rotational strengths arising from the neglect of correlation either in atomic polar tensors or in atomic axial tensors must be comparable, since atomic polar and axial tensors are inextricably linked [7]. The values obtained by calculation in the case of the atomic polar tensors and from differences between theory and experiment for the atomic axial tensors are consistent with this expectation.

#### 3. CONCLUSION

In predicting vibrational spectra, a harmonic force field can be defined as accurate when the residual errors in vibrational energies, wave functions, and transition moments are insignificant relative to the contributions of anharmonicity. There is no practical advantage to be gained from improvements beyond this level. Following recent developments in post-SCF methods, it is now practicable to calculate harmonic force fields of this level of accuracy for chiral molecules whose VCD spectra can be observed experimentally. We have demonstrated this advance here for the specific case of trans-2,3dideuteriooxirane (1). The harmonic force field is obtained from a large basis set MP2 calculation, using semidirect analytical derivative methods. The excellent accuracy of large basis set MP2 harmonic force fields has been documented by Amos, Handy, and coworkers [37-39]. The development [29] of a semidirect implementation of the analytical derivative methodology for the calculation of MP2 harmonic force fields [37,40] permits this accuracy to be realized for oxirane and its isotopomers. The accuracy of the harmonic force field is documented by comparison of calculated and experimental gas-phase vibrational frequencies and unpolarized absorption intensities for oxirane (2) and oxirane-d<sub>4</sub> (3). Errors in vibrational frequencies are attributable predominantly to anharmonicity. Intensities, calculated using MP2 atomic polar tensors [41], are within experimental error with the exception of C-H and C-D stretching modes; the differences here can be attributed to anharmonicity. The harmonic force field is equally successful in accounting for the frequencies and unpolarized absorption intensities of 1 in solution, showing that solvent effects are not significant. Consequently, errors in VCD intensities predicted for 1, with the exception of the C-H and C-D stretching modes, originate predominantly in the absence of correlation in the "local" atomic axial tensors, calculated at the SCF level of approximation. For the modes whose rotational strengths were reported, these correlation errors are of reasonable magnitude and, with one exception, small.

This work constitutes the first comparison to experimental data of a theoretical prediction of a VCD spectrum based on an accurate harmonic force field. Prior predictions of VCD spectra based on ab initio calculations at the SCF level of approximation [8-25], including those for 1 [10,12,17,36], have achieved considerable success in replicating experimental VCD spectra. However, as a result of the substantially lower accuracy of the force fields used in these calculations, a detailed analysis of the contributions of the various possible sources of error has not been possible. As demonstrated here, accessibility of an accurate harmonic force field removes a major source of error from the VCD spectra and greatly simplifies the analysis of the remaining errors. Given the rapidity with which post-SCF ab initio calculations are increasing in feasibility, the range of chiral molecules for which accurate harmonic force fields are accessible and comparable analyses of VCD intensities are possible can be expected to increase substantially in the very near future. This prospect in turn serves to emphasize the importance of additional experimental studies of small, rigid chiral molecules. Despite the efforts of several groups, the number of molecules in this category for which VCD spectra have been measured remains small. Further, it is vital for analyses of the type presented here that the definition of the quantitative accuracy of experimental rotational strengths be addressed more critically. While the measurement of VCD spectra with excellent signal-to-noise ratios at acceptable spectral resolution is now straightforward [42,43], there still remains room for considerable improvement in regard to both the control of artifactual distortions and the absolute calibration of spectra.

Our work demonstrates that the errors in rotational strengths arising from the absence of correlation in the calculation of "local" atomic axial tensors are not insignificant. Consequently, the development and implementation of post-SCF methods for the calculation of atomic axial tensors is a prerequisite to the accurate calculation of rotational strengths at the harmonic level of approximation. It is to be hoped that these will occur in the near future.

#### 4. REFERENCES

- [1] Stephens, P. J. Journal of Physical Chemistry. Vol. 89, p. 748, 1985.
- [2] Stephens, P. J. Journal of Physical Chemistry. Vol. 91, p. 1712, 1987.
- [3] Stephens, P. J. Croatica Chemica Acta. Vol. 62, p. 429, 1989.
- [4] Hehre, W., L. Radom, P. Schleyer, and J. Pople. <u>Ab Initio Molecular Orbital Theory</u>. New York: Wiley, 1986.
- [5] Amos, R. D. Advances In Chemical Physics. Vol. 67, p. 99, 1987.
- [6] Pulay, P. Advances In Chemical Physics. Vol. 67, p. 241, 1987.
- [7] Stephens, P., K. Jalkanen, R. Amos, P. Lazzeretti, and R. Zanasi. <u>Journal of Physical Chemistry</u>. Vol. 94, p.1811, 1990.
- [8] Lowe, M. A., P. J. Stephens, and G. A. Segal. Chemical Physics Letters. Vol. 123, p. 108, 1986.
- [9] Jalkanen, K. J., P. J. Stephens, R. D. Amos, and N. C. Handy. <u>Journal of the American Chemical Society</u>. Vol. 109, p. 7193, 1987.
- [10] Jalkanen, K. J., P. J. Stephens, R. D. Amos, and N. C. Handy. <u>Journal of the American Chemical Society</u>. Vol. 110, p. 2012, 1988.
- [11] Kawiecki, R., F. Devlin, P. Stephens, R. Amos, N. Handy. Chemical Physics Letters. Vol. 145, p. 411, 1988.
- [12] Jalkanen, K., P. Stephens, P. Lazzeretti, and R. Zanasi. <u>Journal of Physical Chemistry</u>. Vol. 93, p. 6583, 1989.
- [13] Jalkanen, K. J., R. W. Kawiecki, P. J. Stephens, and R. D. Amos. <u>Journal of Physical Chemistry</u>. Vol. 94, p. 7040, 1990.
- [14] Bursi, R., F. Devlin, and P. Stephens. <u>Journal of the American Chemical Society</u>. Vol. 112, p. 9430, 1990.
- [15] Kawiecki, R. W., F. J. Devlin, P. J. Stephens, and R. D. Amos. <u>Journal of Physical Chemistry</u>. Vol. 95, p. 9817, 1991.
- [16] Kawiecki, R. W., Ph.D. Thesis, University of Southern California, Los Angeles, CA, 1988.
- [17] Jalkanen, K. J., Ph.D. Thesis, University of Southern California, Los Angeles, CA, 1989.
- [18] Bursi, R., Ph.D. Thesis, University of Southern California, Los Angeles, CA, 1991.
- [19] Amos, R., N. Handy, A. Drake, and P. Palmieri. <u>Journal of Chemical Physics</u>. Vol. 89, p. 7287, 1988.

- [20] Lowe, M. A., and J. S. Alper. Journal of Chemical Physics. Vol. 92, p. 4035, 1988.
- [21] Dothe, H., M. Lowe, and J. Alper. Journal of Physical Chemistry. Vol. 92, p. 6246, 1988.
- [22] Dothe, H., M. Lowe, and J. Alper. Journal of Physical Chemistry. Vol. 93, p. 6632, 1989.
- [23] Malon, P., T. Keiderling, J. Vang, and J. Chickos. <u>Chemical Physics Letters</u>. Vol. 179, p. 282, 1992.
- [24] Malon, P., L. Mickley, K. Sluis, C. Tam, and T. Keiderling. <u>Journal of Physical Chemistry</u>. Vol. 96, p. 10139, 1992.
- [25] Faglioni, F., P. Lazzeretti, M. Malagoli, R. Zanasi, and T. Prosperi. <u>Journal of Physical Chemistry</u>. In press.
- [26] Freedman, T., K. Spencer, N. Ragunathan, L. Nafie, J. Moore, and J. Schwab. <u>Canadian Journal of Chemistry</u>. Vol. 69, p. 1619, 1991. (Note: Figures 1 and 2 were reproduced with permission. Copyright 1991 National Research Council of Canada).
- [27] Nakanaga, T. Journal of Chemical Physics. Vol. 73, p. 5451, 1980.
- [28] Nakanaga, T. Journal of Chemical Physics. Vol. 74, p. 5384, 1981.
- [29] Trucks, G., M. Frisch, M. Head-Gordon, J. Andres, H. Schlegel, A. Salter. <u>Journal of Chemical</u> Physics. To be submitted.
- [30] Frisch, M. J., et al. GAUSSIAN 92, Revision C. GAUSSIAN Inc., Pittsburg, PA, 1992.
- [31] Jalkanen, K. J., P. J. Stephens, R. D. Amos, and N. C. Handy. <u>Journal of Physical Chemistry</u>. Vol. 92, p. 1781, 1988.
- [32] Stephens, P., K. Jalkanen, P. Lazzeretti, and R. Zanasi. <u>Chemical Physics Letters</u>. Vol. 156, p. 509, 1989.
- [33] Amos, R. D., N. C. Handy, K. J. Jalkanen, and P. J. Stephens. <u>Chemical Physics Letters</u>. Vol. 133, p. 21, 1987.
- [34] Amos, R. D., K. J. Jalkanen, P. J. Stephens. Journal of Physical Chemistry. Vol. 92, p. 5571, 1988.
- [35] Amos, R. D., CADPAC 5.0. Cambridge University, Cambridge, U.K., 1992.
- [36] Stephens, P. J., K. J. Jalkanen, R. W. Kawiecki. <u>Journal of the American Chemical Society</u>. Vol. 112, p. 6518, 1990.
- [37] Simandiras, E. D., N. C. Handy, R. D. Amos. Chemical Physics Letters. Vol. 133, p.324, 1987.
- [38] Simandiras, E. D., R. D. Amos, N. C. Handy, T. J. Lee, J. E. Rice, R. B. Remmington, and H. F. Schaefer. <u>Journal of the American Chemical Society</u>. Vol. 110, p. 1388, 1988.

- [39] Simandiras, E. D., J. E. Rice, T. J. Lee, R. D. Amos, and N. C. Handy. <u>Journal of Chemical Physics</u>. Vol. 88, p. 3187, 1988.
- [40] Handy, N. C., R. D. Amos, J. F. Gaw, J. E. Rice, and E. D. Simandiras. Chemical Physics Letters. Vol.120, p. 151, 1985.
- [41] Simaniras, E. D., R. D. Amos, N. C. Handy. Chemical Physics. Vol. 114, p. 9, 1987.
- [42] Nafie, L. A. Advances in Applied Fourier Transform Infrared Spectroscopy. Ch. 3, p. 67, New York: Wiley, 1988.
- [43] Keiderling, T. A. <u>Practical Fourier Transform Infrared Spectroscopy.</u> Ch. 5, p. 203, New York: Academic Press, 1990.

## NO. OF COPIES ORGANIZATION

- 2 DEFENSE TECHNICAL INFO CTR ATTN DTIC DDA 8725 JOHN J KINGMAN RD STE 0944 FT BELVOIR VA 22060-6218
- 1 DIRECTOR
  US ARMY RESEARCH LAB
  ATTN AMSRL OP SD TA
  2800 POWDER MILL RD
  ADELPHI MD 20783-1145
- 3 DIRECTOR
  US ARMY RESEARCH LAB
  ATTN AMSRL OP SD TL
  2800 POWDER MILL RD
  ADELPHI MD 20783-1145
- 1 DIRECTOR
  US ARMY RESEARCH LAB
  ATTN AMSRL OP SD TP
  2800 POWDER MILL RD
  ADELPHI MD 20783-1145

#### ABERDEEN PROVING GROUND

5 DIR USARL ATTN AMSRL OP AP L (305)

#### NO. OF NO. OF COPIES ORGANIZATION COPIES ORGANIZATION 1 HODA 1 OFFICE OF NAVAL RESEARCH ATTN SARD TT DR F MILTON DEPARTMENT OF THE NAVY PENTAGON ATTN R S MILLER CODE 432 WASHINGTON DC 20310-0103 800 N OUINCY STREET **ARLINGTON VA 22217** 1 HODA ATTN SARD TT MR J APPEL 1 COMMANDER PENTAGON NAVAL AIR SYSTEMS COMMAND WASHINGTON DC 20310-0103 ATTN J RAMNARACE AIR 54111C **WASHINGTON DC 20360** HODA OASA RDA 2 ATTN DR C H CHURCH COMMANDER PENTAGON ROOM 3E486 NAVAL SURFACE WARFARE CENTER WASHINGTON DC 20310-0103 ATTN R BERNECKER R 13 G B WILMOT R 16 COMMANDER SILVER SPRING MD 20903-5000 US ARMY RESEARCH OFFICE ATTN R GHIRARDELLI 5 COMMANDER D MANN NAVAL RESEARCH LABORATORY R SINGLETON ATTN M C LIN R SHAW J MCDONALD P O BOX 12211 **E ORAN** RSCH TRNGLE PK NC 27709-2211 J SHNUR R J DOYLE CODE 6110 DIRECTOR 1 **WASHINGTON DC 20375** ARMY RESEARCH OFFICE **COMMANDER** ATTN AMXRO RT IP LIB SERVICES 2 NAVAL WEAPONS CENTER P O BOX 12211 RSCH TRNGLE PK NC 27709-2211 ATTN T BOGGS CODE 388 T PARR CODE 3895 2 COMMANDER CHINA LAKE CA 93555-6001 US ARMY ARDEC ATTN SMCAR AEE B D S DOWNS SUPERINTENDENT 1 PCTNY ARSNL NJ 07806-5000 NAVAL POSTGRADUATE SCHOOL **DEPT OF AERONAUTICS** 2 **COMMANDER** ATTN D W NETZER US ARMY ARDEC **MONTEREY CA 93940** ATTN SMCAR AEE J A LANNON PCTNY ARSNL NJ 07806-5000 3 AL LSCF ATTN R CORLEY 1 **COMMANDER** R GEISLER US ARMY ARDEC J LEVINE ATTN SMCAR AEE BR L HARRIS EDWARDS AFB CA 93523-5000 PCTNY ARSNL NJ 07806-5000 1 **AFOSR** COMMANDER ATTN J M TISHKOFF US ARMY MISSILE COMMAND **BOLLING AIR FORCE BASE** ATTN AMSMI RD PR E A R MAYKUT **WASHINGTON DC 20332** AMSMI RD PR P R BETTS

REDSTONE ARSENAL AL

NO. OF COPIES	ORGANIZATION	NO. OF COPIES	ORGANIZATION
1	OSD SDIO IST ATTN L CAVENY PENTAGON WASHINGTON DC 20301-7100	3	DIRECTOR SANDIA NATIONAL LABORATORIES DIVISION 8354 ATTN S JOHNSTON P MATTERN D STEPHENSON
1	COMMANDANT USAFAS ATTN ATSF TSM CN FORT SILL OK 73503-5600	1	LIVERMORE CA 94550  BRIGHAM YOUNG UNIVERSITY DEPT OF CHEMICAL ENGINEERING
I	UNIV OF DAYTON RSCH INSTITUTE ATTN D CAMPBELL AL PAP EDWARDS AFB CA 93523	1	ATTN M W BECKSTEAD PROVO UT 84058  CALIFORNIA INSTITUTE OF TECH
1	NASA LANGLEY RESEARCH CENTER ATTN G B NORTHAM MS 168. LANGLEY STATION	•	JET PROPULSION LABORATORY ATTN L STRAND MS 125 224 4800 OAK GROVE DRIVE PASADENA CA 91109
4	HAMPTON VA 23365  NATIONAL BUREAU OF STANDARDS US DEPARTMENT OF COMMERCE	1	CALIFORNIA INSTITUTE OF TECHNOLOGY ATTN F E C CULICK MC 301 46 204 KARMAN LAB PASADENA CA 91125
	ATTN J HASTIE M JACOX T KASHIWAGI H SEMERJIAN WASHINGTON DC 20234	1	UNIVERSITY OF CALIFORNIA LOS ALAMOS SCIENTIFIC LAB P O BOX 1663 MAIL STOP B216 LOS ALAMOS NM 87545
2	DIRECTOR LAWRENCE LIVERMORE NATIONAL LAB ATTN C WESTBROOK W TAO MS L 282 P O BOX 808 LIVERMORE CA 94550	1	UNIVERSITY OF CALIFORNIA BERKELEY CHEMISTRY DEPARMENT ATTN C BRADLEY MOORE 211 LEWIS HALL BERKELEY CA 94720
1	DIRECTOR LOS ALAMOS NATIONAL LAB ATTN B NICHOLS T7 MS B284 P O BOX 1663 LOS ALAMOS NM 87545	2	UNIVERSITY OF CALIFORNIA SAN DIEGO ATTN F A WILLIAMS AMES B010 LA JOLLA CA 92093 UNIV OF CALIFORNIA SANTA BARBARA
2	PRINCETON COMBUSTION RESEARCH LABORATORIES INC ATTN N A MESSINA M SUMMERFIELD	2	QUANTUM INSTITUTE ATTN K SCHOFIELD M STEINBERG SANTA BARBARA CA 93106
	PRINCETON CORPORATE PLAZA BLDG IV SUITE 119 11 DEERPARK DRIVE MONMOUTH JUNCTION NJ 08852	1	UNIV OF COLORADO AT BOULDER ENGINEERING CENTER ATTN J DAILY CAMPUS BOX 427 BOULDER CO 80309-0427

NO. OF	ODG ANITZ I TION	NO. OF	ODG ANTI ATTOM
COPIES	ORGANIZATION	COPIES	ORGANIZATION
3	UNIV OF SOUTHERN CALIFORNIA DEPT OF CHEMISTRY ATTN R BEAUDET S BENSON C WITTIG LOS ANGELES CA 90007	4	PENNSYLVANIA STATE UNIVERSITY DEPT OF MECHANICAL ENGINEERING ATTN K KUO M MICCI S THYNELL V YANG UNIVERSITY PARK PA 16802
1	CORNELL UNIVERSITY DEPARTMENT OF CHEMISTRY ATTN T A COOL BAKER LABORATORY ITHACA NY 14853	2	PRINCETON UNIVERSITY FORRESTAL CAMPUS LIBRARY ATTN K BREZINSKY I GLASSMAN P O BOX 710
1	UNIVERSITY OF DELAWARE CHEMISTRY DEPARTMENT ATTN T BRILL NEWARK DE 19711 UNIVERSITY OF FLORIDA	1	PRINCETON NJ 08540  PURDUE UNIVERSITY SCHL OF AERONAUTICS & ASTRONAUTICS ATTN J R OSBORN GRISSOM HALL
-	DEPT OF CHEMISTRY ATTN J WINEFORDNER GAINESVILLE FL 32611	1	WEST LAFAYETTE IN 47906  PURDUE UNIVERSITY  DEPARTMENT OF CHEMISTRY
3	GEORGIA INSTITUTE OF TECHNOLOGY SCHOOL OF AEROSPACE ENGINEERING ATTN E PRICE W C STRAHLE B T ZINN ATLANTA GA 30332	2	ATTN E GRANT WEST LAFAYETTE IN 47906  PURDUE UNIVERSITY SCHL OF MECHANICAL ENGNRNG ATTN N M LAURENDEAU S N B MURTHY
1	UNIVERSITY OF ILLINOIS DEPT OF MECH ENG ATTN H KRIER 144MEB 1206 W GREEN ST URBANA IL 61801	1	TSPC CHAFFEE HALL WEST LAFAYETTE IN 47906  RENSSELAER POLYTECHNIC INST DEPT OF CHEMICAL ENGINEERING ATTN A FONTIJN
1	THE JOHNS HOPKINS UNIV CPIA ATTN T W CHRISTIAN 10630 LITTLE PATUXENT PKWY SUITE 202 COLUMBIA MD 21044-3200	1	TROY NY 12181  STANFORD UNIVERSITY DEPT OF MECHANICAL ENGINEERING ATTN R HANSON STANFORD CA 94305
1	UNIVERSITY OF MICHIGAN GAS DYNAMICS LAB ATTN G M FAETH AEROSPACE ENGINEERING BLDG ANN ARBOR MI 48109-2140	1	UNIVERSITY OF TEXAS DEPT OF CHEMISTRY ATTN W GARDINER AUSTIN TX 78712
1	UNIVERSITY OF MINNESOTA DEPT OF MECHANICAL ENGINEERING ATTN E FLETCHER MINNEAPOLIS MN 55455	1	VA POLYTECH INST AND STATE UNIV ATTN J A SCHETZ BLACKSBURG VA 24061

NO. OF COPIES	ORGANIZATION	NO. OF COPIES	ORGANIZATION
1	APPLIED COMBUSTION TECHNOLOGY INC ATTN A M VARNEY P O BOX 607885 ORLANDO FL 32860	1	HERCULES INC ATTN R V CARTWRIGHT 100 HOWARD BLVD KENVIL NJ 07847
2	APPLIED MECHANICS REVIEWS ASME ATTN R E WHITE & A B WENZEL 345 E 47TH STREET NEW YORK NY 10017	1	ALLIANT TECHSYSTEMS INC MARINE SYSTEMS GROUP ATTN D E BRODEN MS MN50 2000 600 2ND STREET NE HOPKINS MN 55343
1	TEXTRON DEFENSE SYSTEMS ATTN A PATRICK 2385 REVERE BEACH PARKWAY EVERETT MA 02149-5900 BATTELLE	1	ALLIANT TECHSYSTEMS INC ATTN R E TOMPKINS MN 11 2720 600 SECOND ST NORTH HOPKINS MN 55343
1	TWSTIAC 505 KING AVENUE COLUMBUS OH 43201-2693	1	IBM CORPORATION RESEARCH DIVISION ATTN A C TAM 5600 COTTLE ROAD
1	COHEN PROFESSIONAL SERVICES ATTN N S COHEN 141 CHANNING STREET REDLANDS CA 92373	1	SAN JOSE CA 95193  IIT RESEARCH INSTITUTE ATTN R F REMALY 10 WEST 35TH STREET
1	EXXON RESEARCH & ENG CO ATTN A DEAN ROUTE 22E ANNANDALE NJ 08801	1 .	CHICAGO IL 60616  LOCKHEED MISSILES & SPACE CO ATTN GEORGE LO
1	GENERAL APPLIED SCIENCE LABS INC 77 RAYNOR AVENUE RONKONKAMA NY 11779-6649		3251 HANOVER STREET DEPT 52 35 B204 2 PALO ALTO CA 94304
1	GENERAL ELECTRIC ORDNANCE SYSTEMS ATTN J MANDZY 100 PLASTICS AVENUE PITTSFIELD MA 01203	1	OLIN ORDNANCE ATTN V MCDONALD LIBRARY P O BOX 222 ST MARKS FL 32355-0222
1	GENERAL MOTORS RSCH LABS PHYSICAL CHEMISTRY DEPARTMENT ATTN T SLOANE WARREN MI 48090-9055	1	PAUL GOUGH ASSOCIATES INC ATTN P S GOUGH 1048 SOUTH STREET PORTSMOUTH NH 03801-5423
2	HERCULES INC ATTN W B WALKUP E A YOUNT P O BOX 210	1	HUGHES AIRCRAFT COMPANY ATTN T E WARD PO BOX 11337 TUCSON AZ 85734-1337

**ROCKET CENTER WV 26726** 

## NO. OF COPIES ORGANIZATION

- 1 SCIENCE APPLICATIONS INC ATTN R B EDELMAN 23146 CUMORAH CREST WOODLAND HILLS CA 91364
- 3 SRI INTERNATIONAL
  ATTN G SMITH
  D CROSLEY
  D GOLDEN
  333 RAVENSWOOD AVENUE
  MENLO PARK CA 94025
- 1 STEVENS INSTITUTE OF TECH DAVIDSON LABORATORY ATTN R MCALEVY III HOBOKEN NJ 07030
- 1 SVERDRUP TECHNOLOGY INC LERC GROUP ATTN R J LOCKE MS SVR 2 2001 AEROSPACE PARKWAY BROOK PARK OH 44142
- 1 SVERDRUP TECHNOLOGY INC ATTN J DEUR 2001 AEROSPACE PARKWAY BROOK PARK OH 44142
- 3 THIOKOL CORPORATION
  ELKTON DIVISION
  ATTN R BIDDLE
  R. WILLER
  TECH LIB
  P O BOX 241
  ELKTON MD 21921
- 3 THIOKOL CORPORATION
  WASATCH DIVISION
  ATTN S J BENNETT
  P O BOX 524
  BRIGHAM CITY UT 84302
- 1 UNITED TECHNOLOGIES RSCH CENTER ATTN A C ECKBRETH EAST HARTFORD CT 06108
- 1 UNITED TECHNOLOGIES CORP CHEMICAL SYSTEMS DIVISION ATTN R R MILLER P O BOX 49028 SAN JOSE CA 95161-9028

## NO. OF COPIES ORGANIZATION

- 1 UNIVERSAL PROPULSION COMPANY ATTN H J MCSPADDEN 25401 NORTH CENTRAL AVENUE PHOENIX AZ 85027-7837
- 1 VERITAY TECHNOLOGY INC ATTN E B FISHER 4845 MILLERSPORT HIGHWAY EAST AMHERST NY 14051-0305
- 1 FREEDMAN ASSOCIATES
  ATTN E FREEDMAN
  2411 DIANA ROAD
  BALTIMORE MD 21209-1525
- 3 ALLIANT TECHSYSTEMS
  ATTN C CANDLAND
  L OSGOOD
  R BECKER
  600 SECOND ST NE
  HOPKINS MN 55343
- 1 US ARMY BENET LABORATORY
  ATTN SAM SOPOK
  SMCAR CCB B
  WATERVLIET NY 12189

#### NO. OF

#### COPIES ORGANIZATION

#### ABERDEEN PROVING GROUND

#### 36 DIR USARL

ATTN: AMSRL-WT-P, A HORST

AMSRL-WT-PC,

R A FIFER

G F ADAMS

W R ANDERSON

R A BEYER

S W BUNTE

C F CHABALOWSKI

K P MCNEILL-BOONSTOPPEL

A COHEN

R CUMPTON

R DANIEL

D DEVYNCK

N F FELL

B E FORCH

J M HEIMERL

A J KOTLAR

M R MANAA

W F MCBRATNEY

K L MCNESBY

K L MCNLOD

S V MEDLIN

M S MILLER

A W MIZIOLEK

S H MODIANO

J B MORRIS

J E NEWBERRY

S A NEWTON

R A PESCE-RODRIGUEZ

**B M RICE** 

R C SAUSA

M A SCHROEDER

J A VANDERHOFF

**M WENSING** 

A WHREN

J M WIDDER

C WILLIAMSON

AMSRL-CI-CA, R PATEL

INTENTIONALLY LEFT BLANK.

#### USER EVALUATION SHEET/CHANGE OF ADDRESS

This Laboratory undertakes a continuing effort to improve the quality of the reports it publishes. Your comments/answers to the items/questions below will aid us in our efforts. 1. ARL Report Number <u>ARL-TR-963</u> Date of Report <u>February 1996</u> 2. Date Report Received 3. Does this report satisfy a need? (Comment on purpose, related project, or other area of interest for which the report will be used.) 4. Specifically, how is the report being used? (Information source, design data, procedure, source of ideas, etc.) 5. Has the information in this report led to any quantitative savings as far as man-hours or dollars saved, operating costs avoided, or efficiencies achieved, etc? If so, please elaborate. 6. General Comments. What do you think should be changed to improve future reports? (Indicate changes to organization, technical content, format, etc.) Organization CURRENT Name **ADDRESS** Street or P.O. Box No. City, State, Zip Code 7. If indicating a Change of Address or Address Correction, please provide the Current or Correct address above and the Old or Incorrect address below. Organization OLD Name **ADDRESS** Street or P.O. Box No. City, State, Zip Code

(Remove this sheet, fold as indicated, tape closed, and mail.)
(DO NOT STAPLE)